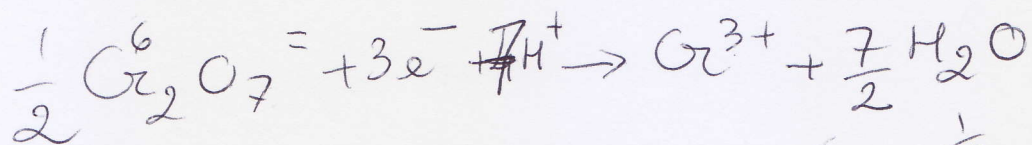


$$E^{\circ} = 1,33$$



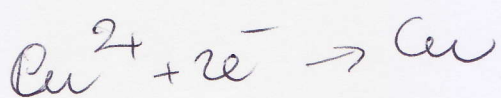
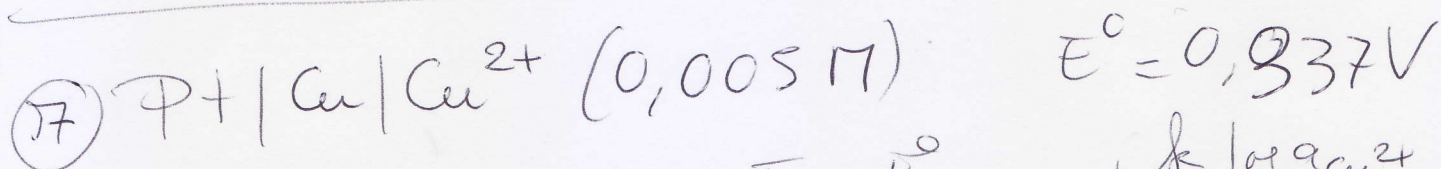
$$E = E^{\circ}_{\text{Cr}_2\text{O}_7^{6-}/\text{Cr}^{3+}, \text{H}^{+}} + \frac{k}{3} \log \frac{(a_{\text{Cr}_2\text{O}_7^{6-}})^{\frac{1}{2}} (a_{\text{H}^{+}})^7}{(a_{\text{Cr}^{3+}}) (a_{\text{H}_2\text{O}})^{7/2}}$$

qui
transcurabile

$$= 1,33 + \frac{0,0591597}{3} \log \frac{(0,01)^{1/2} (10^{-1,5})^7}{(0,001)}$$

$$= 1,162 \text{ V}$$

Redox



$$E = E^{\circ}_{\text{Cu}^{2+}/\text{Cu}} + \frac{k}{2} \log a_{\text{Cu}^{2+}}$$
$$= 0,337 + \frac{0,0591597}{2} \log 0,005$$

$$= 0,269 \text{ V}$$

I specie